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unshared pair. a pair of valence electrons that is not shared

between atoms (lone pair) double covalent bond.

bond involving shared pairs of electrons. triple covalent bond. a bond formed by sharing three pairs of electrons.

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**Coordinate**  
covalent bond. a  
covalent bond in  
which one atom  
contributes both  
bonding electrons.

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### ... Covalent

When H forms a bond with H O to form the

hydronium ion H O , this bond is called a coordinate

covalent bond

because a. both bonding electrons come from the

oxygen atom. b. it forms an especially strong bond. c. the

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electrons are  
equally shared. d.  
the oxygen no  
longer has eight  
valence electrons.

## Key

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...

a pair of valence  
electrons not  
shared between  
atoms. double



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covalent bond. a bond that involves two shared pairs of electrons. triple covalent bond. a bond formed by sharing three pairs of electrons.

coordinate

covalent bond. a covalent bond in which one atom contributes both bonding electrons.

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polyatomic ion.

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Key

-A covalent Bond is the sharing of valence electrons-A molecule is what forms when 2 or more atoms bond covalently o Vs ionic compound-

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Covalent bonds generally occur between elements that are near each other on the periodic table. In particular, between nonmetal elements.

Electronegativity- the attraction that an atom has to a shared pair of electrons.

Elements in the upper right

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corner of the  
periodic table have  
a very strong  
attraction for  
shared pairs of  
electrons-This  
means that they ...

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*bonding.docx -*

*Chem ...*

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### Bonding Crossword

Down : 2) A

covalent bond in which the electrons are shared equally between the two atoms. 3) A bond formed when two atoms share a pair of electrons ( 2 electrons).

*Vocabulary*

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### *Chapter 8 Covalent Bonding ...*

single covalent

bond: atoms joined together by sharing a pair of (2)

electrons: double

covalent bond:

atoms joined

together by sharing

two pairs of (4)

electrons: triple

covalent bond:

atoms joined

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together by sharing  
three pairs of (6)  
electrons:  
molecular

compound:

compound

composed of a  
neutral group of  
atoms joined by  
covalent bonds:

molecular formula

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*"Covalent Bonding"*

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Chapter 8 Covalent Bonding. chemical bond. covalent bond. molecule. electron dot diagram. □□□the force that holds two atoms together. □□□the chemical bond that results from sharing valence electro.... □□a molecule is formed



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When two or more atoms bond covalently.

□□□□ used to show valence electrons of atoms.

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When sharing of electrons occurs, the attachment

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between atoms that results is called a. covalent bond. When such an attachment is formed, bond dissociation energy is released, and the process is. exothermic. When two or more atoms bond by means of electron sharing, the resulting

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particle is called a.  
molecule.

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COVALENT

BONDING Class 8.2

8.2 8.2 8.2 8.3 8.3

8.3 8.3 195

Chapter Quiz

Choose the best  
answer and write  
its letter on the

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line. 1. A bond in which each atom contributes two electrons is a. a double covalent bond. b. an ionic bond. c. a polar covalent bond. d. a coordinate covalent bond. 2. The electron dot structure for hydrogen sulfide,  $H_2S$ , is b.

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Covalent

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Chapter 8 Review

Covalent Bonding

Vocabulary 1.

Covalent bond - a  
bond formed by  
sharing of  
electrons between  
atoms 2. Molecule -  
a neutral group of  
atoms joined

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together by  
covalent bonds 3.

*Thomas Brady  
Covalent Bonding -  
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Single Covalent  
Bonds When only  
one pair of  
electrons is shared,  
such as in a  
hydrogen  
molecule, it is a

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Single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Types of covalent bonds. A. Single covalent bond- 1 shared pair of valence electrons: 2 dots, or a single dash, represent 2 electrons that are simultaneously being attracted by, or "shared" by, the



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nuclei of two neighboring atoms.

1. The formula in the center is a type of structural formula called a “Lewis dot structural formula”.

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242 Chapter 8 •  
Covalent Bonding  
Single Covalent

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Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, Chapter 8: Covalent

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Bonding Based on  
relative  
electronegativities,  
Vocabulary

## Review Answer Key

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